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0.100M potassium sulfate solution is added to a 100.mL solution of 0.200M barium nitrate.

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problems: 1. How many grams of silver chromate will precipitate when 150. mL of 0.500 M silver nitrate are added to 100. mL of 0.400 M potassium chromate? 2  
$$\text{AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{KNO}_3(\text{aq})$$

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Answer Each Of The Following Question.  
Be Sure To Le Complete Sentences Where  
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2.2: Mole to Quantity Stoichiometry

Directions: Solve the following

hypothetical stoichiometry problems.

Assume water is available. 1. When 6.5 mol of solid potassium chlorate breaks into solid potassium chloride and oxygen

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gas, what mass of solid potassium chloride  
is produced? 2.

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1) If I have 340 mL of a 0.5 M NaBr solution, what will the concentration be if I add 560 mL more water to it? 0.19 M (the final volume is 900 mL, set up the equation from that)

2) If I dilute 250 mL of 0.10 M lithium acetate solution to a volume of 750

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mL, what will the concentration of this  
solution be?

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Precipitate If 418.6 ML Of 0.353 M  
Sodium Sulfate Are Allowed To React  
With Excess Barium Nitrate? 2. Consider

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As we learned previously, double replacement reactions involve the reaction between ionic compounds in solution and, in the course of the reaction, the ions in the two reacting compounds are “switched” (they replace each other). Because these reactions occur in aqueous

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solution, we can use the concept of molarity to directly calculate the number of moles of reactants or products that will ...

13.8: Solution Stoichiometry - Chemistry  
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Chemistry I Worksheets And Their Solutions

Stoichiometry is the method that you use



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to figure out how much stuff you'll make in a chemical reaction, or how much stuff you'll need to make a set amount of some product. I'm not going to go into it in huge detail, but I will refer you to a tutorial where I go over the basics in great detail.

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