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CHEMICAL REACTIONS AND EQUATIONS#5 -Completing chemical equation.CLASS X/NTSE/NEET \u0026 IIT FOUNDATION10th Class Chemistry, ch 12, Chemical Reaction of Alkanes Ch 12 Matric Class Chemistry MOLE CoNcept : **STOICHIOMETRY : Class X , XI , XII : CBSE /ICSE**

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We have amount methane and need molecules of hydrogen, so we use molar ratio of moles hydrogen to moles methane. Get the molar mass of methane by adding  $1 \times$  molar mass carbon +  $4 \times$  molar mass of hydrogen. Using the initial amount, divide it by number of grams in 1 mole. Then, multiply that by the molar ratio.

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Stoichiometry Stoichiometry CHAPTER 11 SOLUTIONS MANUAL Section 11.1

Defining Stoichiometry pages 368–372 Practice Problems pages 371–372

1. Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is observed. a.  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$  ...

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