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Covalent  
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Formation.

Chemical Bonding

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Chemistry

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(6th of 19  
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~~Types Of Bonding~~

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11th Class

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Geometry Ionic

Radius Ionic vs.

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Enthalpy FSc

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Higher Ionization



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Ionization Energy

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Atomic Sizes Ch 6

Lec 5 Covalent

Radii. FSc part 1

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Bond: Covalent vs.

Ionic and Polar vs.

Nonpolar FSc

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CH 6, LEC 2:

Causes of Chemical

Bonding FSc

Chemistry Book1,

CH 6, LEC 3:

Energetics of Bond

formation Ch 6 Lec

1 What is Chemical

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Bonding FSc Part 1

Chemistry FSC

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5: Ionic Radii Ch 6

Lec 3 Atomic Radii.

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~~theory 11th Class~~

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CHAPTER 6

REVIEW Chemical

Bonding SECTION

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## Chapter 6

### 1 SHORT ANSWER

Answer the following questions in the space

provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons

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## Chapter 6

(d) Lewis structures 2. b A covalent bond consists of (a) a shared electron.

## Answer Key

6 Chemical Bonding  
- Effingham County  
School District  
Chapter 6: Chemical  
Bonding Section 1-  
Introduction to  
Chemical Bonding  
Objectives: define

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## Chapter 6

Chemical bond;  
differentiate  
between covalent  
and ionic bonding;  
explain why  
bonding occurs; use  
the...

Ch 6 -

HonorsChemWins

Chemical Bonding

sectiOn 1

Introduction to

Chemical Bonding

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## Chapter 6

Section 2 Covalent

Bonding and

Molecular

Compounds section

3 Ionic Bonding and

Ionic Compounds

section 4 Metallic

Bonding section 5

Molecular

Geometry

CHAPTER 6.

Chemistry

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Introduction to  
Chemical Bonding  
Key Terms  
chemical bond  
nonpolar-covalent  
bond ...

Answer Key

### CHAPTER 6

hemical Bonding

A chemical bond is  
a mutual electrical  
attraction between  
the nuclei and  
valence electrons of

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## Chapter 6

different atoms that binds the atoms together. When atoms form a chemical bond, their valence electrons are redistributed to make the atoms more stable. The way the electrons are redistributed determines the type of bond. Chemical bonding that results

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from the electrical

Bonding

CHAPTER 6

Chemical Bonding

Modern Chemistry

1 Chemical Bonding

CHAPTER 6

Chemical Bonding

SECTION 1

Introduction to

Chemical Bonding

OBJECTIVES 1.

Define Chemical

bond. 2. Explain

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## Chapter 6

why most atoms form chemical bonds. 3. Describe ionic and covalent bonding. 4.

## Answer Key

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CHAPTER 6

REVIEW Chemical  
Bonding SECTION  
3 SHORT ANSWER

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Answer the

following questions  
in the space

provided. 1. a The  
notation for sodium

chloride, NaCl,  
stands for one (a)

formula unit. (c)  
crystal. (b)

molecule. (d) atom.

2. d In a crystal of  
an ionic compound,

each cation is  
surrounded by a

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number of (a)  
molecules. (c)  
dipoles. (b) positive  
ions. (d) negative  
ions.

## Answer Key

6 Chemical Bonding

- Somerset

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Chemical Bonds.

Jennie L. Borders.

Standards. SPS1.

Students will

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investigate our  
current  
understanding of  
the atom. b.

Compare and  
contrast ionic and  
covalent bonds in  
terms of electron  
movement. SPS2.  
Students will  
explore the nature  
of matter, its  
classification and its  
system for naming

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types of matter.

Bonding

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Chapter 6 Notes

Chemical Bonding

Section

1:Introduction to

Chemical Bonding

Atoms seldom exist



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Chapter 6

as independent  
particles in nature.  
Most

Section 2

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Chemical Bonding

Section 1 ...

Chapter 6 Review

Section 1 Chemical

Bonding Answers

Modern Chemistry

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Homework 5 6

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Chapter 6

Answers. Semester

1 Exam Study

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51 Chemical. Point

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6 Chemical Bonding  
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more atoms held together by a chemical bond is called what? answer choices. A molecule. A molecule. Tags: Question 37 . SURVEY .

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6 REVIEW

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Chemical Bonding

SECTION 3 SHORT

ANSWER Answer

the following

questions in the

space provided 1 a

The notation for

sodium chloride,

NaCl, stands for

one (a) formula unit

(c) crystal (b)

molecule (d) atom 2

d In a crystal of an

ionic compound,



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Chapter 6

each cation is ...

Bonding

Chapter 6 Chemical

Bonding Section 2

Covalent Answer

Key ...

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SECTION 3 SHORT

ANSWER ANSWER

THE FOLLOWING

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QUESTIONS IN

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THE NOTATION

FOR SODIUM

CHLORIDE  $\text{NaCl}$ '

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Bonding Questions

for Review and

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Chemistry The

Molecular

Chapter 6 Study

*Page 34/38*

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Covalent Answer

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6 Section 3

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## Chapter 6

Chemical Bonding

Chapter 6 Section 3

Chemical  
abouturier.

chemistry chapter 6

section 3,4,5. ionic

compound. formula

unit. lattice energy.

polyatomic ion. A

compound that

consists of positive

and negative ions.

the lowest whole-

number ratio of ions

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Chapter 6

in an ionic  
compound. Page  
3/10

Section 2

Covalent

Answer Key

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