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Coffee Cup

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Heat of a Metal by

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Specific Heat

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Phase Changes:

Exothermic or

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Heat of Metal Sample
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Problem solved

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Lab Answers Pdf

Correct Answer: A.

Substance A. A

chemist mixes 500 g
of lead at $500\text{ }^{\circ}\text{C}$ with
1,200 g of water at
 $20\text{ }^{\circ}\text{C}$. She then mixes
500 g of copper at
 $500\text{ }^{\circ}\text{C}$ with 1,200 g
of water at $20\text{ }^{\circ}\text{C}$.

The specific heat
capacity of lead is
 $0.1276\text{ J/g }^{\circ}\text{C}$ and
the specific heat

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capacity of copper is $0.3845 \text{ J/g}^\circ \text{C}$. What will be true about the final temperatures of the two systems?

Correct Answer: C.

Calorimetry Lab
Flashcards | Quizlet
Investigate how
calorimetry can be
used to find relative
specific heat values

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When different substances are mixed with water. Modify initial mass and temperature values to see effects on the system. One or any combination of the substances can be mixed with water. A dynamic graph (temperature vs. time) shows temperatures of the individual

Access Free Calorimetry Gizmo substances after mixing. Assessment Question

Calorimetry Lab

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Run Gizmo: Test your predictions using the Gizmo. Record your results in the table. ...

Select the MAGNETIC FORCES tab and experiment with the

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Gizmo. Try to answer the questions below. Share your findings with your classmates and teacher. ... You will use the Calorimetry Lab Gizmo™ to determine the specific heat capacities of various ...

Student Exploration:

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Magnetism (ANSWER
KEY)

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Gizmo Warm-up A
calorimeter is an
insulated container
filled with a liquid,
usually water. When a

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A hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance 's specific heat capacity.

Student Exploration:
Seed Germination

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(ANSWER KEY)

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