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~~704 CHAPTER 22 Identify the product that balances the following nuclear reaction: $^{212}_{84}\text{Po} \rightarrow ^{208}_{82}\text{Pb} + ^4_2\text{He} + 1$. The total mass number and atomic number must be equal on both sides of the equation. $^{212}_{84}\text{Po} \rightarrow ^{208}_{82}\text{Pb} + ^4_2\text{He} +$ mass number: $212 - 208 = 4$ atomic number: $84 - 82 = 2$. The nuclide has a mass number of 4 and an atomic number of 2, ^4_2He .~~

~~CHAPTER 22 Nuclear Chemistry~~

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Veritas/ Chemistry/ Chapter 4 Answer Key 3 Practice Problems 13. Use dimensional analysis and the given densities to make the following conversions. a. 14.8 g of boron (B) to cubic centimeters of boron. The density of boron is 2.34 g/ cm³. $(14.8 \text{ g}) (1 \text{ cm}^3 / 2.34 \text{ g}) = 6.32 \text{ cm}^3$ b. 2.8 L of argon (Ar) to grams of argon.

~~Chapter 4 Problem Solving in Chemistry Answer Key~~

1) No, it is not a proper answer; you do not know whether the professor meant homework problem number 20 or 20 homework problems. Twenty is a quantity and does not make sense without a unit. 3) Middle Zeros: Zeros that appear between other nonzero digits are ALWAYS significant.

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